Cambridge Assessment International Education

Cambridge O Level

CHEMISTRY **5070/21**

Paper 2 Theory **May/June 2022**

MARK SCHEME

Maximum Mark: 75

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the May/June 2022 series for most Cambridge IGCSE, Cambridge International A and AS Level and Cambridge Pre-U components, and some Cambridge O Level components.

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Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

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GENERIC MARKING PRINCIPLE 5: Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).
GENERIC MARKING PRINCIPLE 6: Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

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Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	'List rule' guidance For questions that require ***n*** responses (e.g. State **two** reasons …): • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked *ignore* in the mark scheme should not count towards ***n***. • Incorrect responses should not be awarded credit but will still count towards ***n***. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

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6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

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Question	Answer	Marks
1(a)	Na_2CO_3	**1**
1(b)	NH_4Cl	**1**
1(c)	K_2SO_3	**1**
1(d)	$BaSO_4$	**1**
1(e)	Na_3PO_4	**1**
1(f)	KI	**1**

Question	Answer	Marks
2(a)	2, 8, 5	**1**
2(b)	boiling point has a trend / melting point does not have a trend	**1**
2(c)	melting point is below 1000 °C and boiling point is above 1000 °C / 1000 °C is between the melting point and the boiling point	**1**
2(d)(i)	correct dot-and-cross diagram	**1**
2(d)(ii)	weak intermolecular forces / intermolecular forces are easy to break or overcome	**1**
2(e)	**Any two from:** conducts electricity (1) conducts heat (1) malleable (1) ductile (1) sonorous (1) lustrous (1)	**2**

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Question	Answer	Marks
2(f)	moles of nitrogen = 0.6857 **OR** 19.2 / 28 (1) volume = 0.6857 × 24 **OR** 16.457 (1) volume = 16 (dm^3) (1)	**3**

Question	Answer	Marks
3(a)	**A** carbon dioxide (1) **B** water (1) **C** carbon monoxide (1)	**3**
3(b)	butanoic acid (1) $CH_3CH_2CH_2CO_2H$ (1)	**2**
3(c)	butyl ethanoate (1) $CH_3CO_2CH_2CH_2CH_2CH_3$ (1)	**2**

Question	Answer	Marks
4(a)	18	**1**
4(b)	18	**1**
4(c)	$^{40}_{19}K$	**1**
4(d)	loses one electron	**1**

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Question	Answer	Marks
4(e)(i)	lattice (structure) (1) strong attraction between positive and negative ions / strong electrostatic attraction between ions (1)	**2**
4(e)(ii)	**any two from:** conducts electricity as a molten liquid (1) does not conduct electricity as a solid (1) dissolves in water (1)	**2**

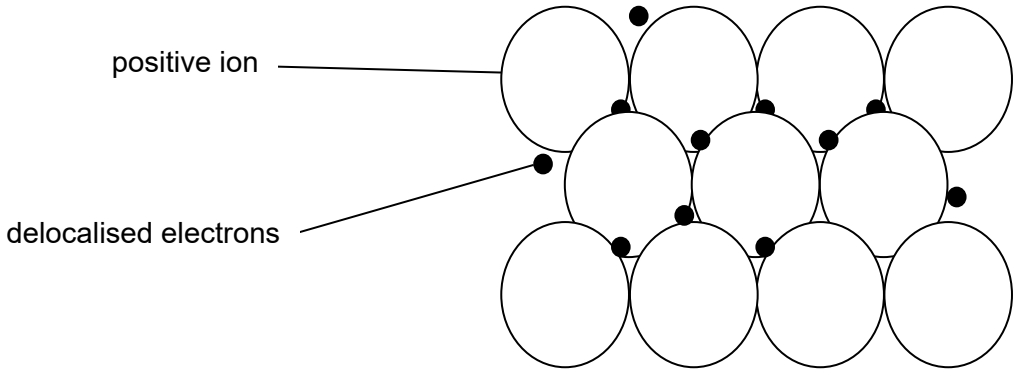
Question	Answer	Marks
5(a)(i)	amount of hydrogen decreases (1) (equilibrium shifts to the left) to absorb energy (1)	**2**
5(a)(ii)	amount of hydrogen stays the same (1) (position of equilibrium does not shift) same number of moles of **gas** on both sides of equation (1)	**2**
5(b)(i)	$Fe(s) + 2H^+(aq) \rightarrow Fe^{2+}(aq) + H_2(g)$ balancing (1) correct state symbol dependent on correct formulae (1)	**2**
5(b)(ii)	add aqueous sodium hydroxide / add aqueous ammonia (1) iron(II) sulfate gives a green precipitate (1)	**2**
5(b)(iii)	278	**1**
5(b)(iv)	13.9	**1**

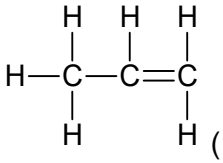
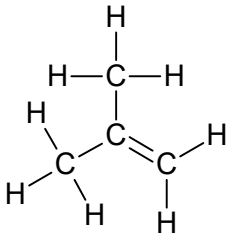
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Question	Answer	Marks
5(c)	cathode – hydrogen / H_2 **AND** anode – oxygen / O_2	**1**
5(d)	iron(III) oxide reacts with carbon to make iron / iron(III) oxide reacts with carbon monoxide to make iron (1) **and any two from:** carbon reacts with oxygen to make carbon dioxide (1) carbon dioxide reacts with carbon to make carbon monoxide (1) calcium carbonate decomposes to make calcium oxide and carbon dioxide (1) calcium oxide reacts with silicon dioxide to make calcium silicate (1)	**3**

Question	Answer	Marks
6(a)	nitrogen and oxygen react together	**1**
6(b)(i)	carbon monoxide gains oxygen so is oxidised (1) and nitrogen monoxide loses oxygen so is reduced (1)	**2**
6(b)(ii)	bond making releases energy and bond breaking absorbs energy (1) more energy released than absorbed (1)	**2**
6(b)(iii)	**particles** move faster / particles have more **kinetic** energy (1) more successful collisions / more collisions above that of activation energy (1)	**2**
6(b)(iv)	lowers the activation energy	**1**
6(b)(v)	more exposed particles (1) more collisions per second (1)	**2**

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Question	Answer	Marks
7(a)	aircraft / trains / spacecraft / ships because it has a low density or electrical cables because of high electrical conductivity / low density (1)	**1**
7(b)	has a layer of oxide / aluminium oxide layer (1) layer is impermeable (to water) / coating is impermeable (to water) (1)	**2**
7(c)	 closely packed positive ions surrounded by delocalised electrons (1) strong attraction between electrons and positive ions (1)	**2**
7(d)	cathode: $Al^{3+} + 3e^- \rightarrow Al$ (1) anode: $2O^{2-} - 4e^- \rightarrow O_2$ / $2O^{2-} \rightarrow O_2 + 4e^-$ (1)	**2**
7(e)	moles of aluminium = 0.0866666667 (1) moles of H_2SO_4 = 0.100 (1) 0.1 mole of acid would react with 0.067 mol of aluminium / 0.0867 mole of aluminium would need 0.13 moles of H_2SO_4 (1)	**3**

Question	Answer	Marks
8(a)	propene	**1**
8(b)	**any** unbranched alkene e.g.  (1) **any** branched alkene e.g.  (1)	**2**
8(c)	bromine (1) alkene goes colourless / decolourised **and** alkane stays orange / does not change colour (1)	**2**
8(d)(i)	**Any two from:** short(er) chain hydrocarbons are in more demand / are more useful (1) provides a source for more gasoline (1) provides a source for hydrogen (1) (feedstock) to make polymers (1)	**2**

Question	Answer	Marks
8(d)(ii)	amount of C is 85.7 / 12 (mol) **and** amount of H is 14.3 / 1 (mol) **OR** mole ratio C : H is 7.14 : 14.3 (1) empirical formula is CH_2 (1) molecular formula C_3H_6 (1)	**3**

Question	Answer	Marks
9(a)	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ balanced equation (1) use of reversible symbol (1)	**2**
9(b)(i)	nitric acid	**1**
9(b)(ii)	M_r is 80 (1) % is 35 (1)	**2**
9(c)(i)	eutrophication	**1**
9(c)(ii)	idea that nitrate fertilisers are very soluble / nitrate fertilisers are more soluble	**1**
9(d)(i)	reduce soil acidity / raise pH of soil	**1**
9(d)(ii)	$2NH_4NO_3 + Ca(OH)_2 \rightarrow Ca(NO_3)_2 + 2NH_3 + 2H_2O$ (1) ammonia is lost / ammonia is formed (1)	**2**

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