Cambridge
O Level

Cambridge International Examinations
Cambridge Ordinary Level

CHEMISTRY **5070/22**

Paper 2 Theory **May/June 2017**

MARK SCHEME

Maximum Mark: 75

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2017 series for most Cambridge IGCSE®, Cambridge International A and AS Level and Cambridge Pre-U components, and some Cambridge O Level components.

® IGCSE is a registered trademark.

© UCLES 2017

Question	Answer	Mark
A1(a)	Copper(II) chloride	**1**
A1(b)	Ammonium chloride	**1**
A1(c)	Hydrogen chloride	**1**
A1(d)	Ammonium chloride	**1**
A1(e)	Carbon tetrachloride	**1**

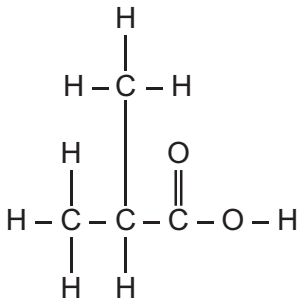
Question	Answer	Mark
A2(a)	<table><tr><th>particle</th><th>atomic number</th><th>number of neutrons in particle</th><th>number of electrons in particle</th></tr><tr><td>^{35}Cl</td><td>17</td><td>18</td><td>**17** (1)</td></tr><tr><td>**^{37}Cl** (1)</td><td>17</td><td>20</td><td>17</td></tr><tr><td>$^{39}K^{+}$</td><td>19</td><td>**20** (1)</td><td>18</td></tr><tr><td>$^{79}Br^{-}$</td><td>**35** (1)</td><td>44</td><td>36</td></tr><tr><td>^{81}Br</td><td>35</td><td>**46** (1)</td><td>35</td></tr><tr><td>**$^{85}Rb^{+}$** (1)</td><td>37</td><td>48</td><td>36</td></tr></table>	**6**
A2(b)(i)	Atoms with same number of protons but different number of neutrons / atoms with same atomic number but different nucleon number / atoms of the same element with different number of neutrons	**1**
A2(b)(ii)	^{35}Cl and ^{37}Cl	**1**

© UCLES 2017

Question	Answer	Mark
A3(a)(i)	Hydrochloric acid **AND** magnesium oxide	**1**
A3(a)(ii)	1 mark each for any 4 of: • Use of **excess** base (1) • Use hot acid / use warm acid / warm the mixture (of acid and base) (1) • Filter mixture (to get filtrate) (1) • Evaporate some of filtrate and allow to crystallise / leave in warm place to crystallise / heat to crystallisation point (1) • (Filter), wash with organic solvent / dry with filter paper / dry in a (drying) oven (1)	**4**
A3(b)	$Ba^{2+}(aq) + SO_4^{2-}(aq) \rightarrow BaSO_4(s)$ Correct formulae and balancing (1) Correct state symbols – dependent on correct formulae (1)	**2**
A3(c)(i)	Moles of acid = 0.020×0.65 **OR** 0.013 (1) Mass = 2.26(2) (g) / 2.3 (g) (1)	**2**
A3(c)(ii)	Percentage yield = 76.(1) %	**1**

Question	Answer	Mark
A4(a)	Sodium ion: 2.8 (1) Oxide ion: 2.8 (1)	**2**
A(b)	Negative electrode: $Na^+ + e^- \rightarrow Na$ (1) Positive electrode: $2O^{2-} \rightarrow O_2 + 4e^-$ (1)	**2**
A(c)	Ions move / mobile ions / ions free to move	**1**
A(d)	$Na_2O + H_2O \rightarrow 2NaOH$	**1**

© UCLES 2017

Question	Answer	Mark
A5(a)	(Acidified) potassium manganate(VII) / oxygen	**1**
A5(b)(i)	Lithium / sodium / potassium / calcium / magnesium (1) Corresponding ethanoate **AND** hydrogen (1)	**2**
A5(b)(ii)	$CaCO_3 + 2CH_3CO_2H \rightarrow Ca(CH_3CO_2)_2 + H_2O + CO_2$ (2) IF: two marks not scored H_2O and CO_2 as products = 1 mark	**2**
A5(c)		**1**
A5(d)(i)	Condensation	**1**
A5(d)(ii)	Decomposes / decays / will not fill up land-fill sites / less litter / no need for incineration	**1**

Question	Answer	Mark
A6(a)	Energy / enthalpy on vertical axis **AND** progress of reaction / course of reaction on horizontal axis (1) Reactant level above product level and to the left of product **AND** reactants and products labelled (1) Enthalpy change shown by downward arrow **AND** labelled enthalpy change or ΔH (1)	**3**
A6(b)	1 mark each for any **two** of: • Lower activation energy (1) • More particles have energy equal to / greater than the activation energy (1) • Different pathway / different mechanism / via an enzyme complex (1) • more successful collisions (between groups on enzyme and substrates) / number of effective collisions increase (with specific groups on enzyme surface) (1)	**2**
A6(c)	Idea that combustion **AND** respiration increase levels of carbon dioxide / carbon in the atmosphere (1) Idea that photosynthesis reduces levels of carbon dioxide / carbon in the atmosphere (1) Idea that these processes balance each other (1)	**3**

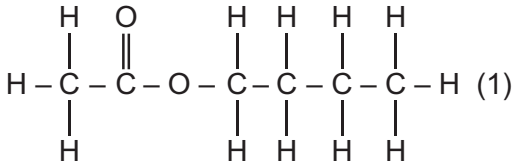
Question	Answer	Mark
B7(a)	Blue solution / bubbles	**1**
B7(b)(i)	Copper(II) sulfate	**1**
B7(b)(ii)	Copper loses electron(s)	**1**
B7(c)	Moles of acid = 0.025×14.0 **OR** 0.35 (1) Moles of sulfur dioxide = 0.175 (1) Volume of gas = 4.2 dm^3 / 4 200 cm^3 (1)	**3**
B7(d)(i)	Blue precipitate / blue solid (which does not redissolve)	**1**
B7(d)(ii)	Blue precipitate / blue solid (1) In excess ammonia gives a dark blue solution (1)	**2**
B7(e)	$2CuCl \rightarrow CuCl_2 + Cu$	**1**

© UCLES 2017

Question	Answer	Mark
B8(a)	Reversible reaction	**1**
B8(b)	(Reaction in which) heat is released	**1**
B8(c)	Idea that no reactants or products can enter / leave	**1**
B8(d)	The colour becomes more brown / colour becomes darker (1) Fewer moles on right hand side so position of equilibrium moves to the left (or reverse argument) / fewer moles on product side so position of equilibrium moves to the left (1)	**2**
B8(e)	The colour more brown / colour becomes darker (1) Exothermic reaction so position of equilibrium moves to the left / backward reaction is endothermic so equilibrium moves to left(1)	**2**
B8(f)(i)	• Add a reactive metal / carbonate to the two acids at the same concentration (1) AND 1 mark for any one of: • Time how long it takes for the metal / carbonate to disappear (1) • Time how long it takes to produce a fixed volume of gas (1) • Count the number of bubbles over fixed time interval (1) • Weak acid has a long**er** reaction time (or reverse argument) / weak acid produces few**er** bubbles in a given time interval (1) **OR** • Add universal indicator to the two acids at the same concentration (1) AND 1 mark for either one of: • Compare colour with colour chart (1) • Red with strong acid AND yellow with weak acid (1) **OR** • Dip pH meter into the two acids at the same concentration (1) AND 1 mark for either one of: • Record pH (1) • pH low**er** for strong acid / pH less for strong acid (than for weak acid) (or reverse argument) (1)	**2**
B8(f)(ii)	KNO_2 **AND** KNO_3	**1**

Question	Answer	Mark
B9(a)	Fuel	**1**
B9(b)	Decomposing vegetation	**1**
B9(c)	Climate change / global warming	**1**
B9(d)(i)	H C Cl 0.040 / 1 0.242 / 12 0.718 / 35.5 **OR** 0.040 mol 0.020 mol 0.020 mol (1) $CH_2C l$ (1)	**2**
B9(d)(ii)	$C_2H_4C l_2$	**1**
B9(e)(i)	The (overall) movement of particles from high concentration to a low concentration / mixing due to (random) movement of particles	**1**
B9(e)(ii)	Particles are moving faster / particles have more kinetic energy	**1**
B9(e)(iii)	**Molecules**/**particles** have different (relative formula) masses / **molecules**/**particles** have different (relative molecular) masses (1) Methane (molecules) move or diffuse faster / butane (molecules) move or diffuse more slowly (1)	**2**

© UCLES 2017

Question	Answer	Mark
B10(a)(i)	$C_nH_{2n+1}OH$ / $C_nH_{2n+2}O$	**1**
B10(a)(ii)	Any value between 154 – 164 (°C) (inclusive of these values)	**1**
B10(b)	(Add) yeast (1) Temperature between 5 and 40 °C / no oxygen present / anaerobic (1) (Fractionally) distil (to get ethanol) (1)	**3**
B10(c)	Butyl ethanoate (1)  (1)	**2**
B10(d)	They get slower / they move less rapidly (when temperature decreases) / molecules slow down (when temperature decreases) / molecules have less kinetic energy (when temperature decreases) (1) They / molecules get closer together (when temperature decreases) (1) They / molecules arranged less randomly / less irregularly (when temperature decreases) (1)	**3**