Cambridge IGCSE

Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CANDIDATE NAME	
CENTRE NUMBER	
CANDIDATE NUMBER	

5703974316

CHEMISTRY **0620/33**

Paper 3 Theory (Core) **October/November 2018**

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use an HB pencil for any diagrams or graphs.
Do not use staples, paper clips, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
Electronic calculators may be used.
A copy of the Periodic Table is printed on page 20.
You may lose marks if you do not show your working or if you do not use appropriate units.

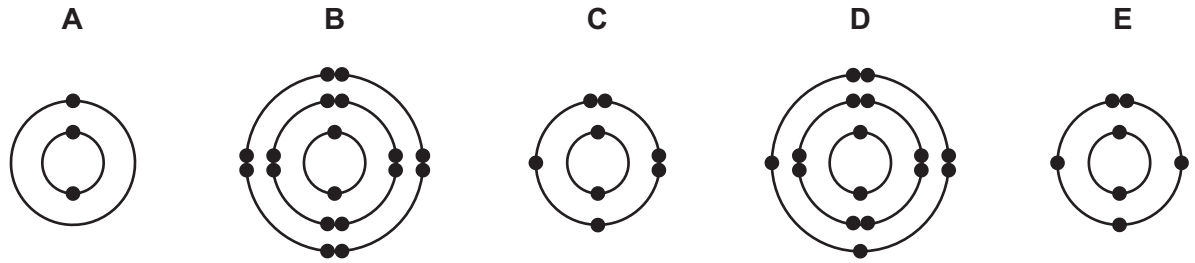
At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **18** printed pages and **2** blank pages.

© UCLES 2018

1 **(a)** The electronic structures of five atoms, **A**, **B**, **C**, **D** and **E**, are shown.



Answer the following questions about these structures.
Each structure may be used once, more than once or not at all.
State which structure, **A**, **B**, **C**, **D** or **E**, represents:

(i) an atom with a total of eight electrons ..... [1]

(ii) an atom in Group V of the Periodic Table ..... [1]

(iii) an atom with a complete outer shell of electrons ..... [1]

(iv) an atom of a metallic element ..... [1]

(v) an atom which forms a stable ion with a single positive charge. ..... [1]

(b) Complete the table to show the number of electrons, neutrons and protons in the neon atom and copper ion shown.

	number of electrons	number of neutrons	number of protons
$^{22}_{10}Ne$	10		
$^{65}_{29}Cu^{+}$		36	

[3]

[Total: 8]

© UCLES 2018

2 **(a)** The table shows the concentrations of the ions present in the solution obtained from squid nerve cells and in human blood plasma.

ion present	solution obtained from squid nerve cells	human blood plasma
	concentration in g / 1000 cm^3	concentration in g / 1000 cm^3
sodium	1.15	3.25
potassium	15.60	0.16
magnesium	0.03	0.04
chloride	3.55	3.65
hydrogencarbonate	trace	1.50

Answer these questions using only information from the table.

(i) Give **two** major differences between the concentrations of the ions present in the solution obtained from squid nerve cells and in human blood plasma.

1 .....

2 .....
[2]

(ii) Calculate the mass of potassium ions present in 250 cm^3 of the solution obtained from squid nerve cells.

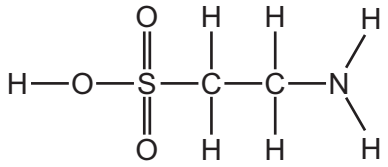
mass of potassium ions = ... g [1]

(b) Describe a test for sodium ions.

test .....

result .....
[2]

(c) Squid produce an ink which contains a compound called taurine.
The structure of taurine is shown.



Deduce the molecular formula of taurine showing the number of carbon, hydrogen, oxygen, nitrogen and sulfur atoms.

..... [1]

(d) The ink used for writing is a mixture of dyes.
These dyes can be separated by paper chromatography.

Describe how to separate a mixture of dyes using paper chromatography.
Include a labelled diagram in your answer.

.....

.....

.....

.....

.....
[4]

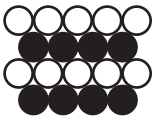
[Total: 10]

© UCLES 2018

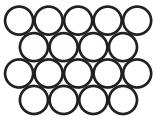
3 **(a)** Brass is an alloy of copper and zinc.

(i) Which **one** of the following diagrams best represents an alloy?

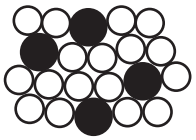
J



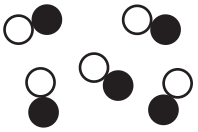
K



L



M



..... [1]

(ii) Brass is used to make the propellers of ships rather than pure copper or pure zinc.

Suggest a property of brass which explains this.

.....

..... [1]

(b) The chemical equation for the reaction of zinc with concentrated nitric acid is shown.

$$Zn + 4HNO_3 \rightarrow Zn(NO_3)_2 + 2NO_2 + 2H_2O$$

(i) Complete the word equation for this reaction.

zinc	+	nitric acid	→	....	+	....	+	.....

[2]

(ii) One of the compounds in this equation is a pollutant gas which contributes to acid rain.

Identify the pollutant gas and state a common source of it.

pollutant gas .....

source .....

[2]

(c) Zinc oxide is reduced by heating it with carbon.

$$ZnO + C \rightarrow Zn + CO$$

How does this equation show that zinc oxide is reduced?

.....

..... [1]

(d) When green iron(II) sulfate is heated it loses its water of crystallisation.
The reaction is reversible.

(i) Complete the following equation by writing the sign for a reversible reaction in the box.

$$\underset{\text{green}}{FeSO_4.7H_2O} \quad \square \quad \underset{\text{white}}{FeSO_4} + 7H_2O$$

[1]

(ii) Use the information in the equation to suggest how to change white iron(II) sulfate into green iron(II) sulfate.

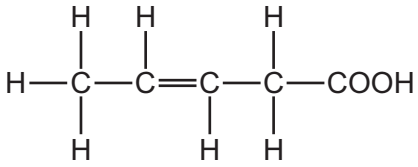
..... [1]

[Total: 9]

© UCLES 2018

4 **(a)** The structure of compound **P** is shown.

compound **P**



(i) What feature of the structure of compound **P** shows that it is unsaturated?

..... [1]

(ii) Describe the colour change when an excess of compound **P** is added to aqueous bromine.

from ..... to .....
[2]

(iii) Compound **P** has a –COOH functional group.

Draw the structure of the –COOH functional group. Show all of the atoms and all of the bonds.

[1]

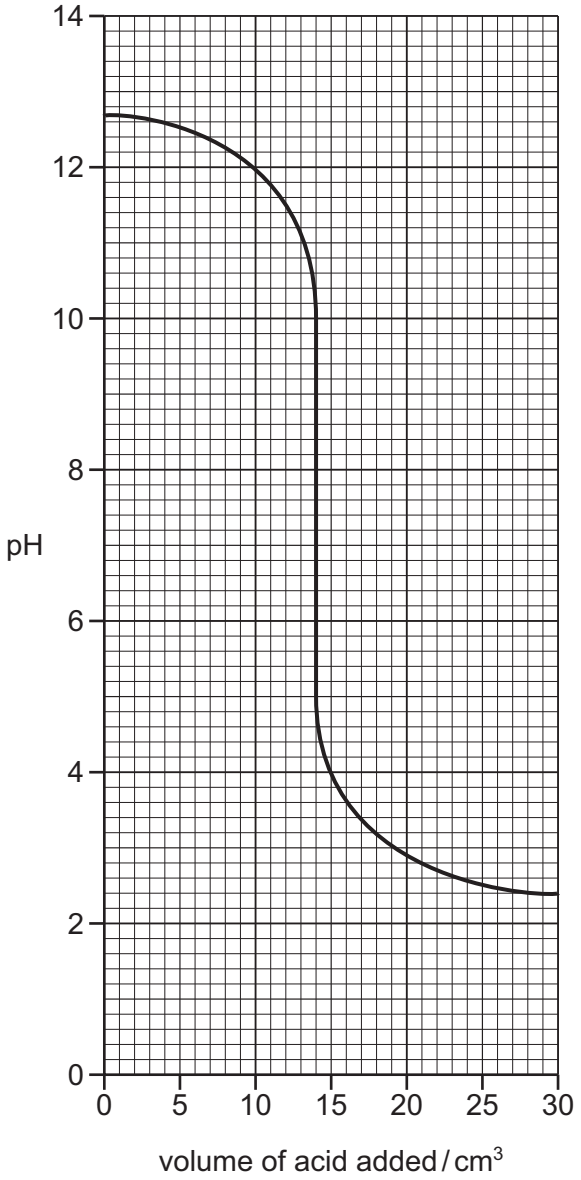
(iv) What effect would compound **P** have on litmus solution?

..... [1]

© UCLES 2018

[Turn over

(b) The concentration of aqueous sodium hydroxide can be found by reacting it with an acid of known concentration.
The graph shows how the pH of aqueous sodium hydroxide in a conical flask changes as acid is added to it.



(i) Describe how the pH changes as the acid is added.

.....

.....

..... [2]

(ii) What is the pH of the aqueous sodium hydroxide before the acid is added?

..... [1]

(iii) What volume of acid has been added when the solution reaches neutral pH?

..... [1]

[Total: 9]

© UCLES 2018

5 **(a)** When sulfur vapour falls on to a cold surface, small crystals of solid sulfur form.

Use the kinetic particle model to describe the arrangement **and** motion of the particles in sulfur when it is:

- a gas

.....

.....

- a solid

.....

.....

[4]

(b) Sulfur dioxide is formed when copper reacts with hot concentrated sulfuric acid.

Balance the chemical equation for this reaction.

$$Cu + ...H_2SO_4 \rightarrow CuSO_4 + SO_2 + ...H_2O$$

[2]

(c) Hot copper reacts with chlorine to form copper(II) chloride.

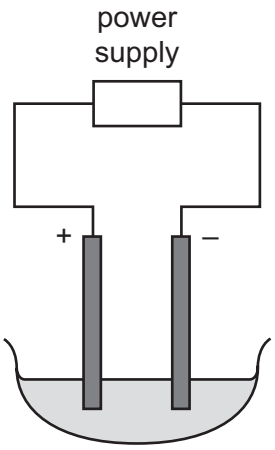
Describe a test for chloride ions.

test .....

result .....

[2]

(d) Molten copper(II) chloride can be electrolysed using the apparatus shown.



(i) On the diagram label:
- the cathode
- the electrolyte

[2]

(ii) Predict the products of this electrolysis at:

the positive electrode .....

the negative electrode. .....

[2]

(iii) Give **one** observation that is made at the positive electrode.

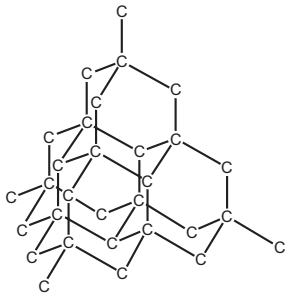
..... [1]

[Total: 13]

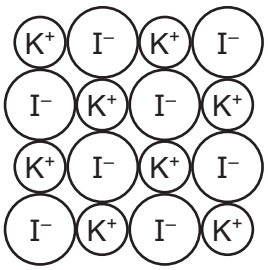
© UCLES 2018

6 **(a)** The diagrams show the structures of four substances, **R**, **S**, **T** and **U**.

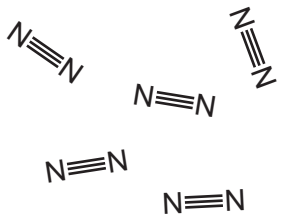
R



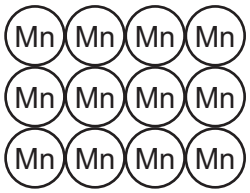
S



T



U



State which **one** of these substances, **R**, **S**, **T** or **U**:

(i) is a compound

..... [1]

(ii) conducts electricity when solid

..... [1]

(iii) has the lowest melting point

..... [1]

(iv) is a macromolecule

..... [1]

(v) can act as a catalyst.

..... [1]

(b) Substance **S** is potassium iodide.
When aqueous potassium iodide reacts with aqueous chlorine, the solution turns brown.

Name the brown substance that forms. Using ideas about the reactivity of the halogens, suggest why the brown substance forms.

.....

..... [2]

(c) Potassium salts are present in many fertilisers.

Which **one** of the following compounds is also present in many fertilisers?
Tick **one** box.

lead(II) bromide	☐
calcium phosphate	☐
copper(II) sulfide	☐
tin(IV) oxide	☐

[1]

(d) Many fertilisers contain ammonium salts.

Explain why farmers do **not** add calcium hydroxide (slaked lime) to the soil at the same time as fertilisers containing ammonium salts.

.....

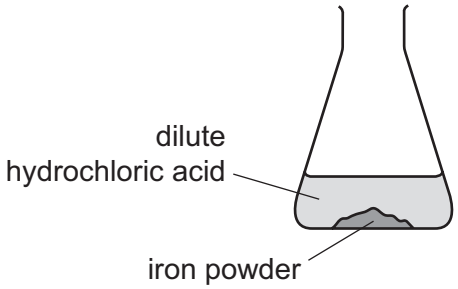
.....

..... [2]

[Total: 10]

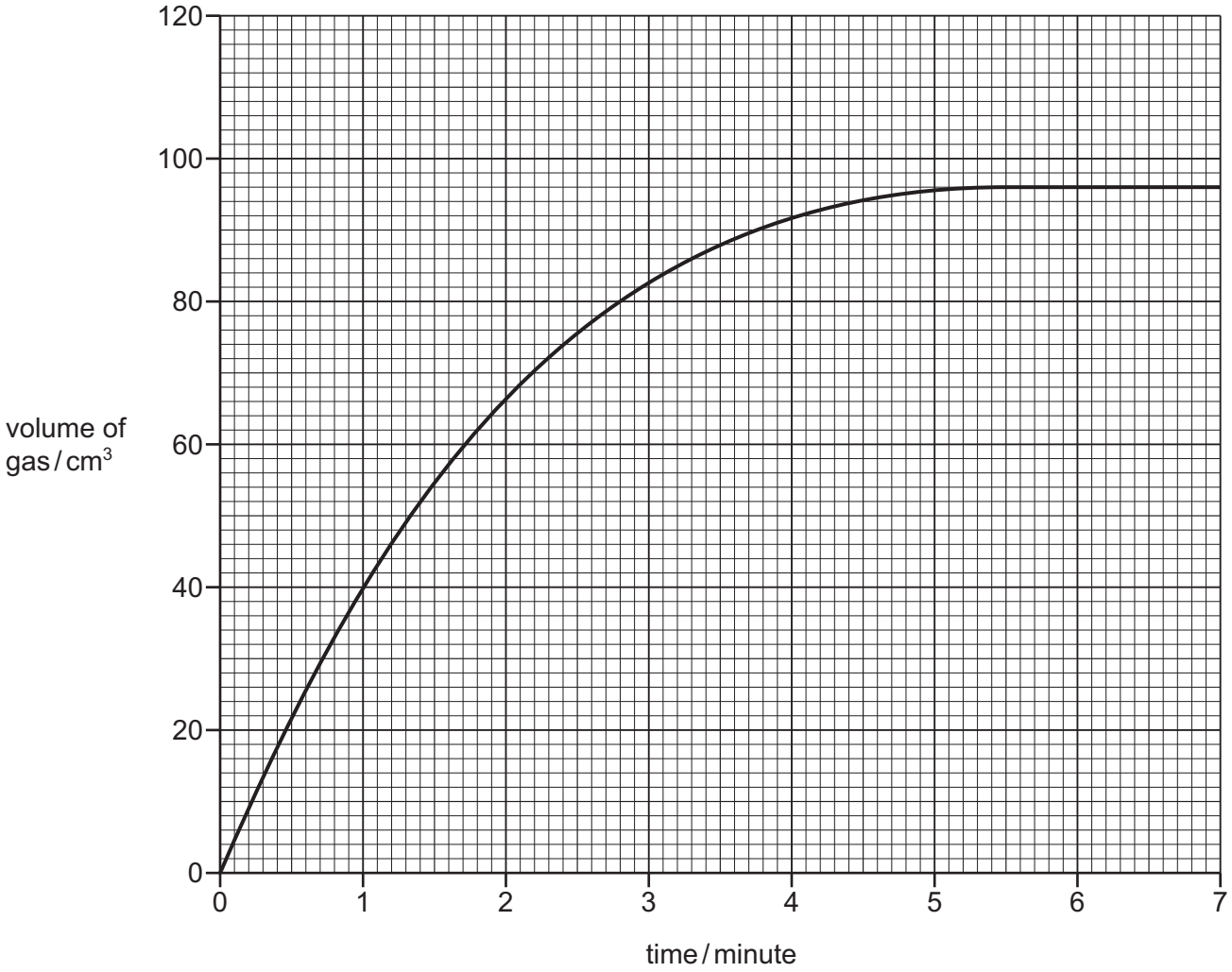
© UCLES 2018

7 A student investigates the rate of reaction of iron powder with dilute hydrochloric acid.



(a) Complete the diagram to show the apparatus the student could use to collect and measure the gas produced.
Label your diagram. [3]

(b) The graph shows the results the student obtained using dilute hydrochloric acid of concentration 0.2 mol / dm^3 and an **excess** of iron powder.



Use the graph to deduce:

(i) the time that the reaction was complete

..... [1]

(ii) the volume of gas produced when the reaction was complete.

..... [1]

(c) The student repeated the experiment using a lower concentration of dilute hydrochloric acid. All other conditions were kept the same.

On the grid, draw a graph to show how the volume of gas changes with time when a lower concentration of dilute hydrochloric acid is used. [2]

© UCLES 2018

(d) Iron is extracted from iron ore using a blast furnace.
The solid substances added to the blast furnace are iron ore, coke and limestone (calcium carbonate).

(i) State the name of an ore of iron.

..... [1]

(ii) Complete the sentences about the reactions which occur in a blast furnace using words from the list.

air **decomposes** **dioxide** **monoxide**

nitrogen **oxidises** **slag** **tetrachloride**

The coke burns in a blast of hot ..... to form carbon dioxide. This reacts with further hot coke to form carbon ..... . This gas reduces the iron(III) oxide in the iron ore to iron.

The limestone ..... to form lime (calcium oxide) which reacts with impurities in the iron to form ..... .

[4]

[Total: 12]

© UCLES 2018

[Turn over

8 The table shows the properties of some Group VII elements.

element	boiling point in °C	density at room temperature in g / cm^3	physical state at room temperature
fluorine	–188	0.0017	
chlorine		0.0032	gas
bromine	59	3.1	liquid
iodine	184	4.9	solid

(a) (i) Use this information to:

- identify the physical state of fluorine at room temperature

.....

- estimate the boiling point of chlorine.

.....
[2]

(ii) Suggest why the density of chlorine is much lower than the densities of bromine and iodine.

.....

..... [1]

(b) Chlorine is used in water treatment.

(i) Why is chlorine added to water?

..... [1]

(ii) State **one** major use of water in industry.

..... [1]

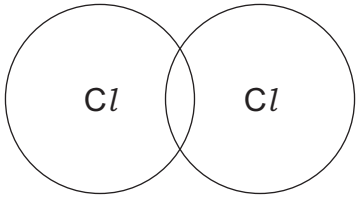
(c) Chlorine reacts with phosphorus to form phosphorus(III) chloride.

Balance the chemical equation for this reaction.

$$\text{.....P} + \text{.....C}l_2 \rightarrow 2\text{PC}l_3$$

[2]

(d) Complete the dot-and-cross diagram to show the electron arrangement in a molecule of chlorine. Show outer shell electrons only.



[2]

[Total: 9]

BLANK PAGE

BLANK PAGE

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

To avoid the issue of disclosure of answer-related information to candidates, all copyright acknowledgements are reproduced online in the Cambridge International Examinations Copyright Acknowledgements Booklet. This is produced for each series of examinations and is freely available to download at www.cie.org.uk after the live examination series.

Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.

© UCLES 2018

The Periodic Table of Elements

Group																	
I	II											III	IV	V	VI	VII	VIII
							1 H hydrogen 1										2 He helium 4
3 Li lithium 7	4 Be beryllium 9											5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20
11 Na sodium 23	12 Mg magnesium 24											13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40
19 K potassium 39	20 Ca calcium 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium –	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131
55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium –	85 At astatine –	86 Rn radon –
87 Fr francium –	88 Ra radium –	89–103 actinoids	104 Rf rutherfordium –	105 Db dubnium –	106 Sg seaborgium –	107 Bh bohrium –	108 Hs hassium –	109 Mt meitnerium –	110 Ds darmstadtium –	111 Rg roentgenium –	112 Cn copernicium –		114 Fl flerovium –		116 Lv livermorium –		

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium –	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium –	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium –	94 Pu plutonium –	95 Am americium –	96 Cm curium –	97 Bk berkelium –	98 Cf californium –	99 Es einsteinium –	100 Fm fermium –	101 Md mendelevium –	102 No nobelium –	103 Lr lawrencium –

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).