Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY **0620/53**

Paper 5 Practical Test **October/November 2016**

MARK SCHEME

Maximum Mark: 40

Published

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Question	Answer	Mark
1(a)	table of results for Experiment 1 temperature boxes completed correctly results comparable to supervisor's	 **1** **1**
1(b)	table of results for Experiment 2 temperature boxes completed correctly results comparable to supervisor's	 **1** **1**
1(c)	all points correctly plotted ± half a small square smooth line graphs labelled	**2** **1** **1**
1(d)(i)	value from graph – 60 s	**1** **1**
1(d)(ii)	value from graph shown clearly	**1** **1**
1(e)	room temperature or initial temperature from results table reaction has finished / stopped	**1** **1**
1(f)	more readings / points / data smoother curve / better or more accurate graph	**1** **1**
1(g)	polystyrene is an insulator / copper is a (good) conductor reduced heat losses	**1** **1**

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Question	Answer	Mark
2(a)(i)	pH 0–3	**1**
2(a)(ii)	effervescence / bubbles / fizzes lighted splint 'pops'	**1** **1**
2(a)(iii)	effervescence / bubbles / fizzes limewater turns milky	**1** **1**
2(a)(iv)	white precipitate	**1**
2(b)(i)	pH 10–14	**1**
2(b)(ii)	white precipitate insoluble / no change	**1** **1**
2(b)(iii)	brown precipitate	**1** **1**
2(b)(iv)	green precipitate	**1**
2(c)	sulfuric acid	**1** **1**
2(d)	calcium hydroxide	**1** **1**

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Question	Answer	Mark
3	**silica** filter (the cleaner) wash the residue dry the residue **water** heat (the filtrate / cleaner) condense the vapour **sodium carbonate** heat to dryness / no liquid left (then solid) sodium carbonate is left **OR** heat until saturated then cool to crystallise / leave to crystallise	**6**

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